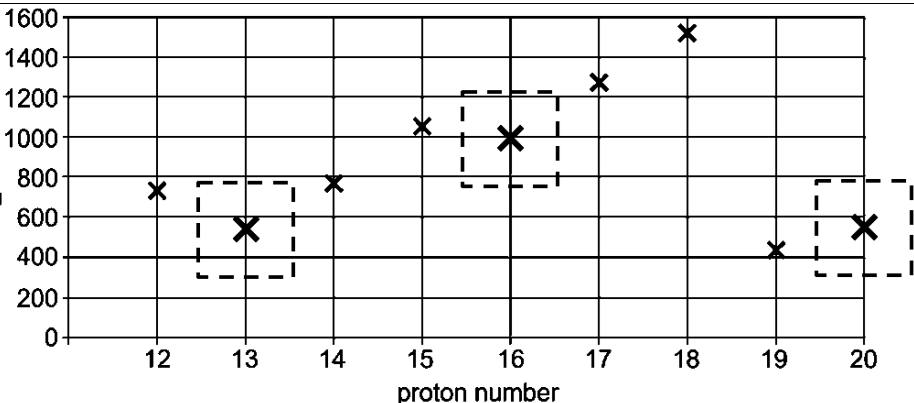
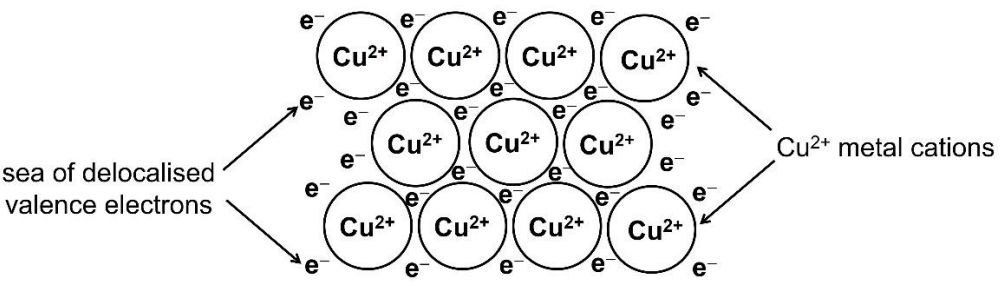


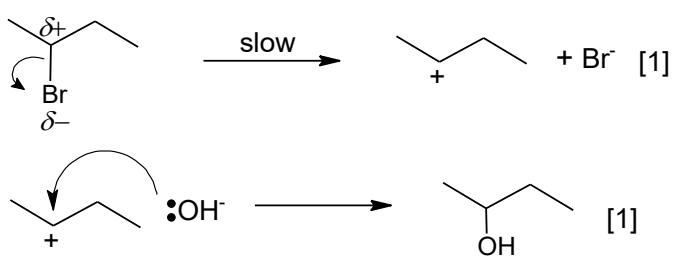
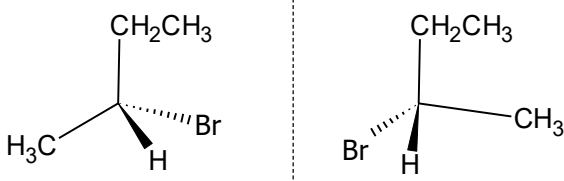

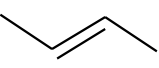
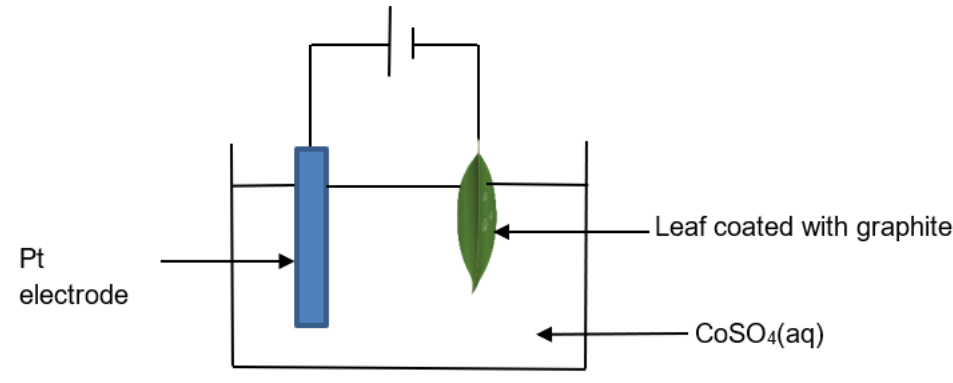
Paper 1 Answer Key

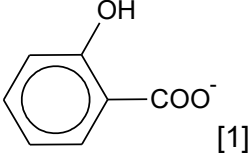
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2	B	12	A	22	D
3	A	13	D	23	C
4	D	14	A	24	C
5	B	15	C	25	D
6	C	16	A	26	C
7	D	17	B	27	C
8	B	18	A	28	D
9	A	19	C	29	D
10	B	20	B	30	A

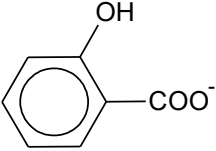
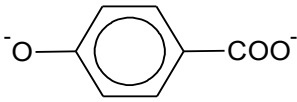
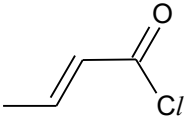
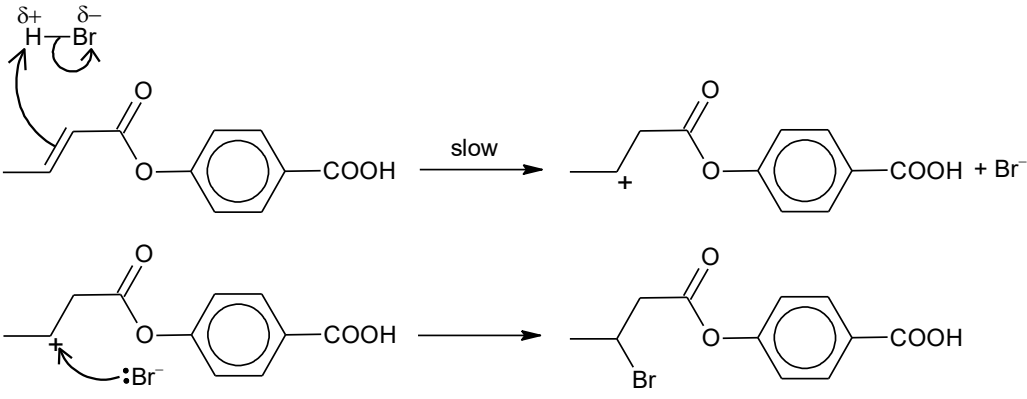
Paper 2 Answers

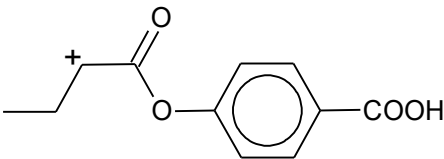
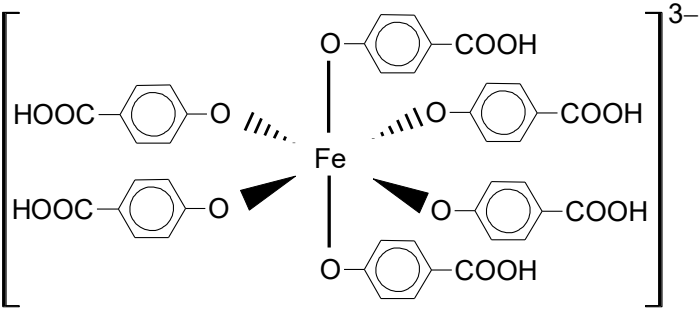
1 (a) (i)	 <p>[1] for all 3 plots and ensure that the plot at proton number 20 is lower than proton number 12.</p>
(ii)	<ul style="list-style-type: none"> • The number of protons increase and nuclear charge increase. • Successive electrons are added to the same valence shell, shielding effect remains approximately constant. • Effective nuclear charge increases and electrostatic attraction between the nucleus and valence electron increases. • More energy is required to remove the valence electron
(iii)	<p>Proton number 15: $1s^2 2s^2 2p^6 3s^2 3p_x^1 3p_y^1 3p_z^1$ Proton number 16: $1s^2 2s^2 2p^6 3s^2 3p_x^2 3p_y^1 3p_z^1$ [1] for both electronic configurations. Reject students who write $1s^2 2s^2 2p^6 3s^2 3p_x^4$ for proton number 16.</p> <p>Element with proton number 16, the valence electron to be removed is from the paired electrons in $3p_x$ orbital. Hence, it experiences inter-electronic repulsion arising from the two electrons occupying the same $3p_x$ orbital and requires less energy to remove. [1]</p>

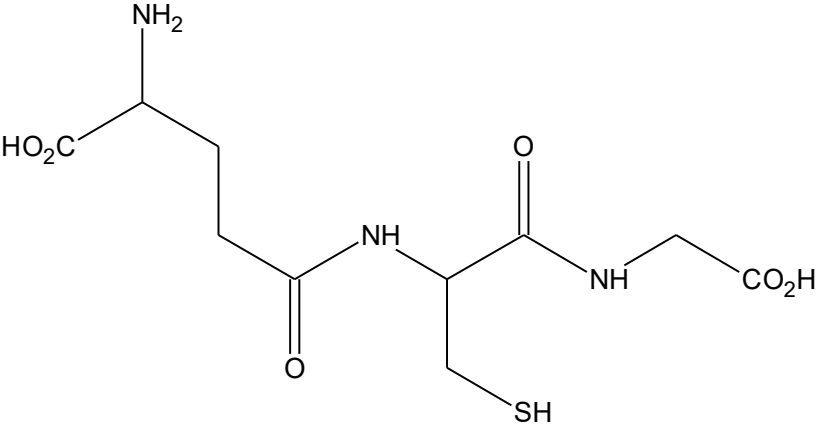
(b) (i)	 <p style="text-align: right;">[1]</p> <p>Copper has a giant metallic structure with strong metallic bonding (or electrostatic attraction) between copper ion and sea of delocalised electrons. The presence of sea of delocalised electrons acts as mobile charge carriers, hence making copper a good conductor of electricity. [1]</p>
(ii)	<ul style="list-style-type: none"> Both Na₂O and MgO have giant ionic structure held together by strong ionic bonding (or electrostatic attraction) between positively charged metal cations and negatively charged oxide anions. $LE \propto \frac{ q^+q^- }{r_+ + r_-}$ Both Na₂O and MgO have the same oxide anion, but Mg²⁺ is doubly charged while Na⁺ is singly charged. The ionic radius of Na⁺ > Mg²⁺. The magnitude of LE is larger for MgO than Na₂O. Ionic bonding of MgO is stronger than Na₂O. More energy is required to break the stronger ionic bonds between Mg²⁺ and O²⁻ in MgO than ionic bonds between Na⁺ and O²⁻ in Na₂O. Melting point of MgO is higher than Na₂O and hence able to withstand higher temperature in case of fire.
2 (a) (i)	<p>The order of reaction with respect to a given reactant is the power to which the concentration of that reactant is raised in the experimentally determined rate equation. [1]</p>
(ii)	<p>When [NaOH] decreases, gradient remains constant, hence rate of reaction is constant. Rate of reaction is independent of [NaOH], hence the reaction is zero order wrt [NaOH]. [1]</p> <p>When [CH₃CHBrCH₂CH₃] = 0.25 mol dm⁻³ initial rate = 0.0048 / 30 = 1.60 × 10⁻⁴ mol dm⁻³ min⁻¹</p> <p>When [CH₃CHBrCH₂CH₃] = 0.50 mol dm⁻³ initial rate = 0.0095 / 30 = 3.16 × 10⁻⁴ mol dm⁻³ min⁻¹</p> <p>When [CH₃CHBrCH₂CH₃] doubles (0.50/0.25), the gradient doubles, showing that the rate of reaction doubles from 1.60 × 10⁻⁴ to 3.16 × 10⁻⁴ mol dm⁻³ min⁻¹. [1]</p> <p>Hence the reaction is first order wrt [CH₃CHBrCH₂CH₃].</p>
(iii)	<p>rate = k [CH₃CHBrCH₂CH₃] 1.60 × 10⁻⁴ = k (0.25) k = 6.40 × 10⁻⁴ min⁻¹ or 1.07 × 10⁻⁵ s⁻¹ [1]</p>

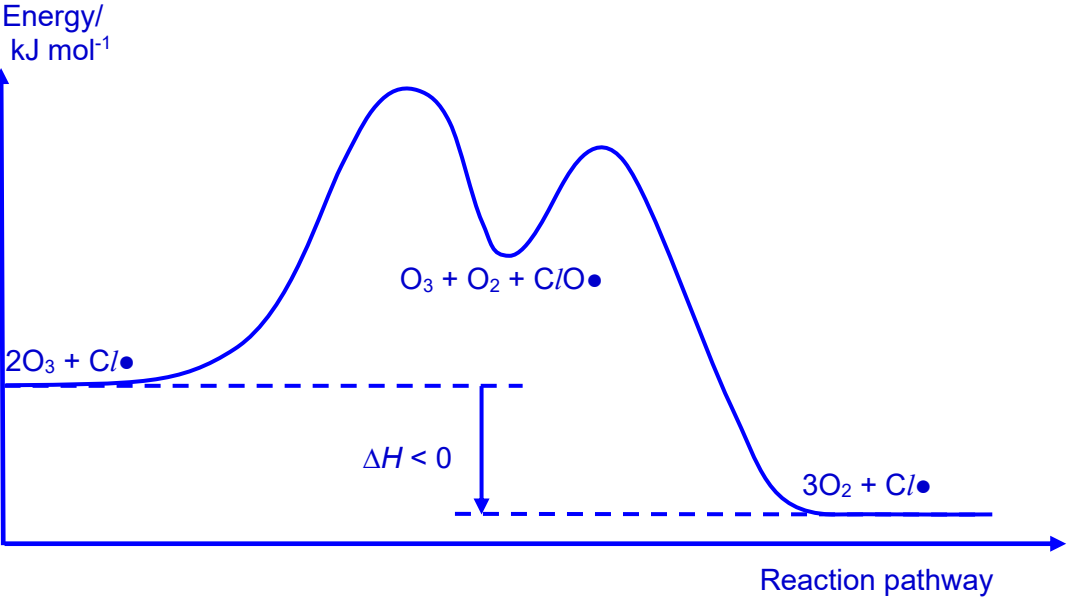
(iv)	<p>Nucleophilic Substitution (S_N1)</p>  <p>[1]</p>
(v)	<p>The OH^- nucleophile can attack the trigonal planar carbocation centre from either the top or bottom of the plane with equal probability, producing equal amounts of the pair of enantiomers (a racemic mixture). [1]</p> <p>The product mixture is optically inactive (i.e. does not rotate plane-polarised light) as each enantiomer rotate plane-polarised light in the opposite direction by the same magnitude hence the rotating powers of the enantiomers cancel out. [1]</p>
(b) (i)	<p>Stereoisomerism is defined as compounds having the same molecular formula but different spatial formulae / different spatial arrangement of atoms. [1]</p>
(ii)	 <p>mirror plane</p> <p>[1]</p>
(iii)	 <p>[1]</p>
(iv)	
3 (a)	 <p>Pt electrode</p> <p>Leaf coated with graphite</p> <p>$\text{CoSO}_4(\text{aq})$</p> <p>[1]</p>

(b)	Cathode: $\text{Co}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Co}(\text{s})$ [1] Anode: $\text{H}_2\text{O}(\text{l}) \rightarrow \text{O}_2(\text{g}) + 4\text{H}^+(\text{aq}) + 4\text{e}^-$ [1]
(c)	$n(\text{Co}) \text{ coated} = (8.90 \times 0.2 \times 10^{-1} \times 10) / 58.9 = 0.0302 \text{ mol}$ [1] $n(\text{O}_2) = (0.0302) / 2 = 0.0151 \text{ mol}$ Volume of gas = $0.0151 \times 24 = 0.363 \text{ dm}^3$ (or 363 cm^3) [1]
(d)	$\text{Co}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Co}(\text{s})$ $Q = It = n z F$ $0.5 \times t = 0.0302 \times 2 \times 96500$ $t = 11657.2\text{s} = 3.24 \text{ h}$ [1] unit must be in h
(e)	$\text{Co}^{2+} + 2\text{e}^- \rightleftharpoons \text{Co} -0.28 \text{ V}$ $\text{Pb}^{2+} + 2\text{e}^- \rightleftharpoons \text{Pb} -0.13 \text{ V}$ $\text{O}_2 + 4\text{H}^+ + 4\text{e}^- \rightleftharpoons 2\text{H}_2\text{O} +1.23 \text{ V}$ The mass of the lead electrode will decrease. $E^\circ(\text{Pb}^{2+}/\text{Pb})$ is more negative than $E^\circ(\text{O}_2/\text{H}_2\text{O})$, Pb will be preferentially oxidised to form Pb^{2+} in aqueous solution. [1] Dull gray solid Pb will be deposited at the leaf. $E^\circ(\text{Pb}^{2+}/\text{Pb})$ is less negative than $E^\circ(\text{Co}^{2+}/\text{Co})$, Pb^{2+} will be preferentially reduced. [1]
4 (a) (i)	Acidity depends on the stability of the conjugate base. In phenoxide, the p-orbital of oxygen atom overlaps with the π electron cloud of the benzene ring. Lone pair of electrons of the oxygen atom delocalised into the benzene ring. The negative charge on the oxygen atom is dispersed into the benzene ring and stabilising the phenoxide. [1] In carboxylate ion, the p-orbital of oxygen atom overlaps with the π electron cloud of the C=O bond. Lone pair of electrons of the oxygen atom delocalised into the C=O bond. The negative charge on the oxygen atom is dispersed over the carbon atom and two electronegative oxygen atoms and stabilising the carboxylate ion. The charge dispersion is more effective than that in phenoxide due to the two electronegative oxygen atoms and hence carboxylate ion is more stable than phenoxide ion. [1] Hence, carboxylic acid is a stronger acid than phenol.
(ii)	 <p>[1]</p> <p>Reject $^-\text{O}-\text{C}_6\text{H}_5-\text{COOH}$ as the COOH should deprotonate first since it has a lower pK_a. Reject $^-\text{O}-\text{C}_6\text{H}_5-\text{COO}^-$ as it differs from 2-hydroxybenzoic acid by 2 H^+ hence not a conjugate base of the acid.</p> <p>The conjugate base of 2-hydroxybenzoic acid can form intramolecular hydrogen bonding that helps to further stabilize the conjugate base. [1]</p>

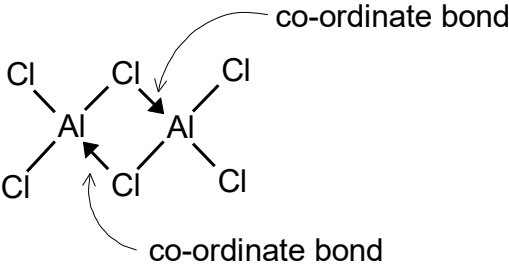
(iii)	<p>$[4\text{-hydroxybenzoic acid}] = \frac{2.76}{138.0} \div \frac{100}{1000} = 0.2000 \text{ mol dm}^{-3}$ [1]</p> <p>$[\text{H}_3\text{O}^+] = \sqrt{10^{-4.58} \times 0.2000} = 2.293 \times 10^{-3} \text{ mol dm}^{-3}$</p> <p>$\text{pH} = -\lg(2.293 \times 10^{-3}) = 2.639 = 2.64$ [1]</p>
(iv)	<p>$n_{4\text{-hydroxybenzoic acid}} = \frac{2.76}{138} = 0.02000 \text{ mol}$</p> <p>$n_{\text{NaOH}} = \frac{0.35}{40.0} = 8.750 \times 10^{-3} \text{ mol}$ [1]</p> <p>Since the neutralisation between 4-hydroxybenzoic acid and NaOH is 1:1 mole ratio, NaOH is a limiting reagent.</p> <p>$n_{4\text{-hydroxybenzoic acid left}} = 0.02000 - (8.750 \times 10^{-3}) = 0.01125 \text{ mol}$</p> <p>$n_{4\text{-hydroxybenzoate formed}} = 8.750 \times 10^{-3} \text{ mol}$ [1]</p> <p>Since there is a mixture of 4-hydroxybenzoic acid and 4-hydroxybenzoate, solution Q is an acidic buffer.</p> <p>$\text{pH} = \text{p}K_a + \lg \frac{[4\text{-hydroxybenzoate}]}{[4\text{-hydroxybenzoic acid}]} = 4.58 + \lg \frac{8.750 \times 10^{-3} / \frac{100}{1000}}{0.01125 / \frac{100}{1000}} = 4.470 = 4.47$ [1]</p>
(v)	<div style="display: flex; justify-content: space-around;"> <div style="border: 1px solid black; padding: 10px; text-align: center;">  <p>2-hydroxybenzoic acid at pH 12</p> <p>position on Fig. 4.1: 3</p> </div> <div style="border: 1px solid black; padding: 10px; text-align: center;">  <p>4-hydroxybenzoic acid at pH 12</p> <p>position on Fig. 4.1: 4</p> </div> </div> <p>[1] for each correct structure and position</p>
(vi)	Any pH between 4.58 and 9.51. [1]
(b) (i)	 <p>[1]</p>
(ii)	<p>Electrophilic Addition</p>  <p>[1] for step 1</p> <p>[1] for step 2</p>

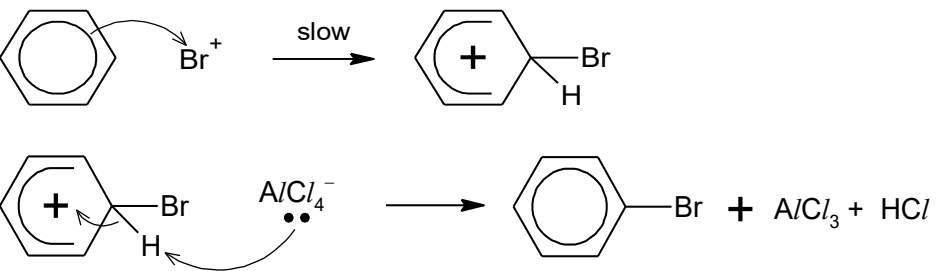
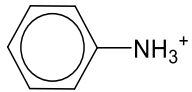
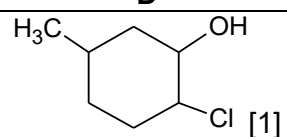
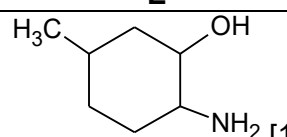
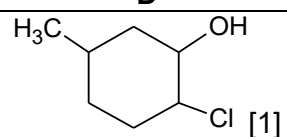
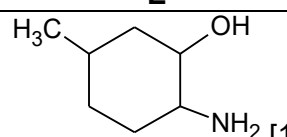
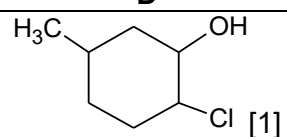
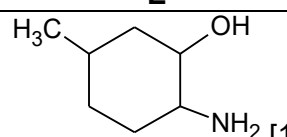
(iii)	<p>Formation of T involves a less stable carbocation that is adjacent to an electron withdrawing ester group, which destabilises the carbocation. Hence, a smaller amount of that carbocation intermediate is formed, and T is a minor product. [1]</p>  <p>(T forms a less stable carbocation than S)</p>
(c) (i)	Fe(OH) ₃ [1]
(ii)	<p>A complex is a species that contains a central metal atom/ion surrounded by molecules or anions (known as ligands) which form coordinate bonds (dative bonds) to the metal centre. [1]</p>
(iii)	<p>The phenoxide ion is a stronger ligand as the lone pair on the O atom is less delocalised compared to that in the carboxylate ion, and hence more available to form dative bonds. OR In the carboxylate ion, the lone pair of electrons of the O atom is delocalised into the C=O bond more effectively due to the electronegative oxygen atom, hence it is less available to form dative bonds. [1]</p>
(iv)	 <ul style="list-style-type: none"> • correct shape • overall charge (overall charge = +3 + 6 (– 1) = –3) • correct structure
5 (a) (i)	<p>When NaOH is added, OH[–] reacts with H⁺, decreasing [H⁺]. By Le Chatelier's Principle, the position of equilibrium for the dissociation of H₂S shifts to the right to increase [H⁺], resulting in higher [S^{2–}]. [1] Precipitation of metal sulfides occurs as the ionic product of metal sulfides becomes greater than the K_{sp}. [1]</p>
(ii)	<p>Volume of 100 g of lake water sample = 100 g / 1.02 g cm^{–3} = 98.03 cm³ = 0.09803 dm³</p> <p>[Cu²⁺] in g dm^{–3} = 6.54 / 0.09803 = 66.71 g dm^{–3} [Cu²⁺] in mol dm^{–3} = 66.71 / 63.5 = 1.050 = 1.05 mol dm^{–3} [1]</p> <p>[Zn²⁺] in g dm^{–3} = 1.25 / 0.09803 = 12.75 g dm^{–3} [Zn²⁺] in mol dm^{–3} = 12.75 / 65.4 = 0.1949 = 0.195 mol dm^{–3} [1]</p>
(iii)	K _{sp} for ZnS = [Zn ²⁺][S ^{2–}] mol ² dm ^{–6} [1]

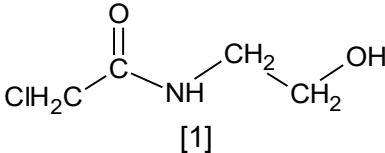
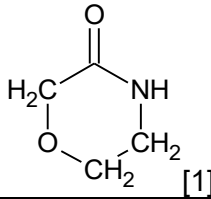
(iv)	<p>For precipitation of metal to occur, ionic product $> K_{sp}$ $[M^{2+}][S^{2-}] > K_{sp}$</p> <p>Minimum $[S^{2-}]$ required for precipitation of CuS = $6.3 \times 10^{-36} / 1.050$ $= 6.00 \times 10^{-36} \text{ mol dm}^{-3}$ [1] allow ecf</p> <p>Minimum $[S^{2-}]$ required for precipitation of ZnS = $1.6 \times 10^{-24} / 0.1949$ $= 8.21 \times 10^{-24} \text{ mol dm}^{-3}$ [1] allow ecf</p> <p>Black precipitate of CuS would appear first, followed by appearance of white precipitate of ZnS. [1]</p>
(v)	<p>$\text{Cu}(\text{NO}_3)_2(\text{s}) + \text{aq} \rightarrow \text{Cu}^{2+}(\text{aq}) + 2\text{NO}_3^{-}(\text{aq}) \text{ -----(1)}$ $\text{CuS}(\text{s}) + \text{aq} \rightleftharpoons \text{Cu}^{2+}(\text{aq}) + \text{S}^{2-}(\text{aq}) \text{ -----(2)}$</p> <p>On addition of solid $\text{Cu}(\text{NO}_3)_2$, concentration of the common ion Cu^{2+} increases (1). By Le Chatelier's Principle, position of equilibrium of dissolving of $\text{CuS}(\text{s})$ (2) shifts to the left to reduce the added Cu^{2+}. Hence, solubility of CuS will decrease and student's suggestion is right. [1]</p> <p>However, solubility product of CuS will remain unchanged since it is affected only by temperature changes and so student's suggestion is wrong. [1]</p>
(b) (i)	<p>SiO_2 has <u>a giant molecular structure</u> [1] where each Si atom is covalently bonded to 4 other O atoms and each O atom is covalently bonded to 2 other Si atoms. [1]</p>
(ii)	<p>The energy released during the formation of instantaneous dipole-induced dipole attraction between water and the Si or O atoms [1] is not enough to overcome the energy needed to break strong covalent bonds between Si and O atoms. [1]</p>
(c) (i)	<p>$2\text{GSH} \rightarrow \text{GSSG} + 2\text{H}^+ + 2\text{e}^-$ [1] Oxidation. GSH loses electron to transform into GSSG. [1]</p>
(ii)	 <p style="text-align: right;">[1]</p>
(d) (i)	<p>High temperature. High kinetic energy results in faster motion of halogen molecules. Hence intermolecular forces of attraction between halogen molecules become less significant. [1]</p> <p>Low pressure. Volume of container is large. Volume occupied by halogen molecules becomes negligible as compared to the volume of container. [1]</p>

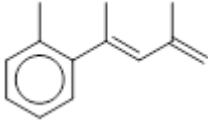
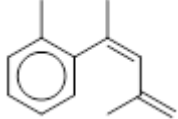
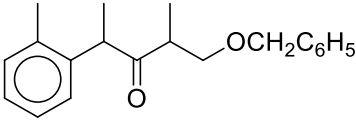
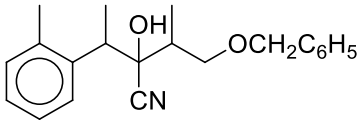
(ii)	C/• acting as a catalyst. It is used in step 1 and regenerated in step 2. [1]
(iii)	 <p>[1] correct shape of diagram showing 2 steps and overall exothermic reaction [1] label axis, <u>balanced</u> chemical species, and enthalpy change correctly</p> <p>Note: 1st step is likely to be endothermic as more energy was taken in to break a stronger O→O bond in ozone (O=O→O) than energy released to form a weaker Cl-O bond. 2nd step has to be exothermic in order for overall ΔH to be exothermic. Optional for first peak to be higher than second peak i.e. E_{a1} > E_{a2}</p>
(e) (i)	<p>[R] $I_2 + 2e \rightarrow 2I^-$ [O] $2S_2O_3^{2-} \rightarrow S_4O_6^{2-} + 2e$ Overall equation: $I_2 + 2S_2O_3^{2-} \rightarrow S_4O_6^{2-} + 2I^-$ $E^\ominus_{\text{cell}} = +0.54 - (+0.09) = +0.45 \text{ V}$ [1] balanced equation and E^\ominus_{cell} value.</p>
(ii)	<p>[R] $Cl_2 + 2e \rightarrow 2Cl^-$ [O] $S_2O_3^{2-} + 5H_2O \rightarrow 2SO_4^{2-} + 10H^+ + 8e$ [1] correct half equations Overall: $4Cl_2 + S_2O_3^{2-} + 5H_2O \rightarrow 8Cl^- + 2SO_4^{2-} + 10H^+$ [1] correct balanced overall equation</p>

Paper 3 Answers

1 (a) (i)	<p>NaCl has a <u>giant ionic structure</u>, it undergoes <u>hydration</u> to form a neutral solution of hydrated ions with <u>pH = 7</u>. $\text{NaCl(s)} + \text{aq} \rightarrow \text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq})$</p> <p>$\text{AlCl}_3$ has a <u>simple covalent structure with ionic character</u> thus it undergoes <u>hydration</u> to give the hydrated ions. $\text{AlCl}_3(\text{s}) + 6\text{H}_2\text{O(l)} \rightarrow [\text{Al}(\text{H}_2\text{O})_6]^{3+}(\text{aq}) + 3\text{Cl}^-(\text{aq})$</p> <p>$[\text{Al}(\text{H}_2\text{O})_6]^{3+}$ then undergoes <u>further hydrolysis</u> to give $[\text{Al}(\text{H}_2\text{O})_5\text{OH}]^{2+}$. Al^{3+} in $[\text{Al}(\text{H}_2\text{O})_6]^{3+}$ has a <u>high charge density</u> and is highly polarising. It is able to <u>polarise one of the O–H bond</u> in H_2O, causing <u>H^+ to be released</u> to give a <u>slightly acidic solution</u> of <u>pH = 3</u>. $[\text{Al}(\text{H}_2\text{O})_6]^{3+}(\text{aq}) + \text{H}_2\text{O(l)} \rightleftharpoons [\text{Al}(\text{H}_2\text{O})_5\text{OH}]^{2+}(\text{aq}) + \text{H}_3\text{O}^+(\text{aq})$</p> <p>$\text{PCl}_5$ has a <u>simple covalent structure</u> and undergoes <u>hydrolysis</u> to give a highly acidic solution of <u>pH = 1</u> and white fumes of HCl gas. $\text{PCl}_5(\text{s}) + 4\text{H}_2\text{O(l)} \rightarrow \text{H}_3\text{PO}_4(\text{aq}) + 5\text{HCl(g)}$</p> <p>[1] for each point</p>									
(ii)	<ul style="list-style-type: none">pH value of BeCl_2: Below 3Quote the ionic radius of Al^{3+} and Be^{2+}: <table><tr><th>Ion</th><th>Ionic radius / nm</th><th>Charge density</th></tr><tr><td>Al^{3+}</td><td>0.050</td><td>$3/0.050 = 60$</td></tr><tr><td>Be^{2+}</td><td>0.031</td><td>$2/0.031 = 64.5$</td></tr></table> <ul style="list-style-type: none">Charge density of Be^{2+} is higher than Al^{3+}; its degree of hydrolysis is higher than Al^{3+}, hence, the pH of BeCl_2 lower than AlCl_3.	Ion	Ionic radius / nm	Charge density	Al^{3+}	0.050	$3/0.050 = 60$	Be^{2+}	0.031	$2/0.031 = 64.5$
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(iii)	 <p>co-ordinate bond</p> <p>co-ordinate bond</p>									
(b)	<p>Both BeO and Al_2O_3 are amphoteric and reacts with both acids and bases. BeO and Al_2O_3 react similarly as Be and Al have similar effective nuclear charge. [1]</p> <p>$\text{BeO(s)} + 2\text{HCl(aq)} \rightarrow \text{BeCl}_2(\text{aq}) + \text{H}_2\text{O(l)}$ $\text{BeO(s)} + 2\text{NaOH(aq)} + \text{H}_2\text{O(l)} \rightarrow \text{Na}_2\text{Be(OH)}_4(\text{aq})$ [1]</p>									
(c) (i)	<p>Br is less electronegative than Cl, hence will react with AlCl_3 to form the electrophile Br^+ which undergoes electrophilic substitution with benzene.</p>									

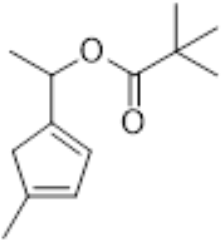
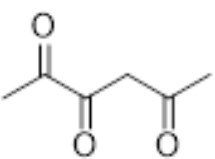
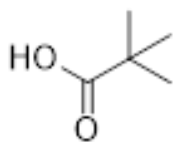
(ii)	<p>Electrophilic substitution</p> $\text{BrCl} + \text{AlCl}_3 \rightleftharpoons \text{Br}^+ + \text{AlCl}_4^-$  <p>3 correct steps – 2 marks 2 correct steps – 1 mark</p>				
(iii)	<ul style="list-style-type: none"> Least: Isomer 3-position Most: Isomer 4-position Electron donating $-\text{C}(\text{CH}_3)_3$ group is directly bonded to the carbocation centre hence it disperses the positive charge on the carbocation centre in the intermediates for isomer 2-position and isomer 4-position more effectively than for isomer 3-position, forming a more stable carbocation. Hence, isomer 3-position will be formed the least. However, the incoming Br^+/electrophile will experience steric hindrance from the bulky $-\text{C}(\text{CH}_3)_3$ group as it approaches the 2-position. Hence, isomer 4-position will be formed the most. 				
(iv)	<ul style="list-style-type: none"> Under acidic conditions, phenylamine undergoes acid-base reaction to produce a high proportion of . The $-\text{NH}_3^+$ is 3-directing and hence a large proportion of the 3-isomer is produced. 				
(d)	<p>B, A, C</p> <ul style="list-style-type: none"> In B, the p orbital of C/ atom overlaps with π electron cloud of benzene ring resulting in the delocalisation of lone pair of electrons on the C/ atom into the benzene ring, strengthening the C–C/ bond, and is least susceptible to attack by water nucleophile. [1] In A, C atom in C–C/ is attached to only one electronegative atom, C/. It is less electron deficient / carries a smaller partial positive charge δ^+ than the carbon atom in –COC/ and is less susceptible to attack by water nucleophile. [1] In C, the C atom in –COC/ is attached to two highly electronegative atoms, O and C/. It is more electron deficient / carries a larger partial positive charge δ^+. This C atom is most highly susceptible to attack by water nucleophile. [1] 				
(e) (i)	<table border="1" data-bbox="359 1758 1332 1937"> <thead> <tr> <th data-bbox="359 1758 845 1792">D</th> <th data-bbox="845 1758 1332 1792">E</th> </tr> </thead> <tbody> <tr> <td data-bbox="359 1792 845 1937">  </td> <td data-bbox="845 1792 1332 1937">  </td> </tr> </tbody> </table>	D	E		
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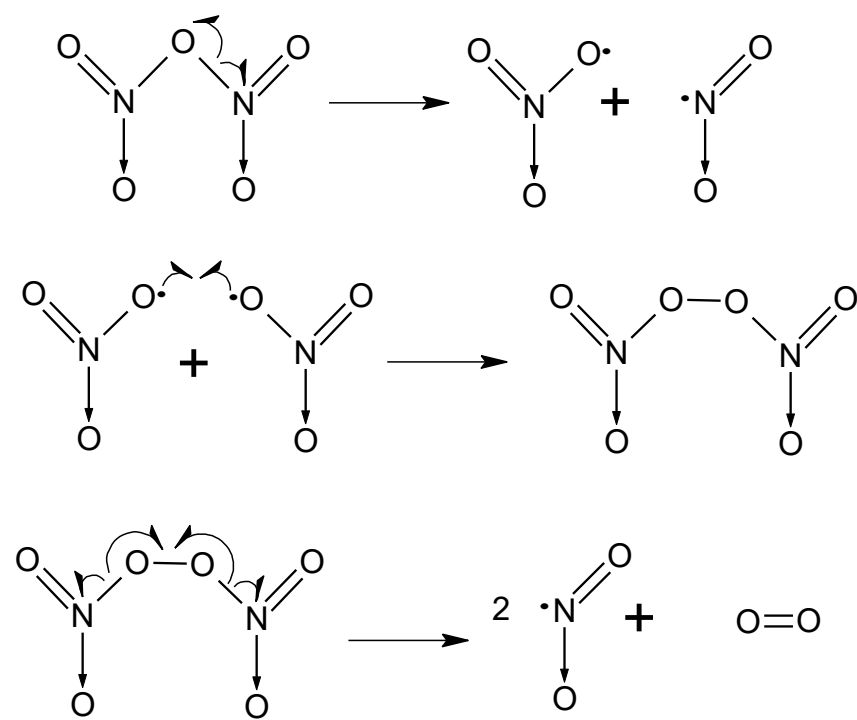
(ii)	<p style="text-align: center;">F</p>  <p style="text-align: center;">[1]</p>	<p style="text-align: center;">G</p>  <p style="text-align: center;">[1]</p>
2 (a)	<p>heat energy from heating coil = heat absorbed by water + heat absorbed by calorimeter</p> $10000 = (300 \times 4.18 \times 7.5) + (C \times 7.5) \quad [1]$ $10000 = (9405) + (7.5C)$ $7.5C = 10000 - 9405 = 595 \text{ J}$ <p>Heat capacity of calorimeter, $C = \frac{595}{7.5} = 79.3 \text{ J K}^{-1} \quad [1]$</p>	
(b) (i)	$K_{sp} = [\text{Ag}^+]^2[\text{CO}_3^{2-}] = (2.57 \times 10^{-4})^2(2.57 \times 10^{-4}/2) = 8.49 \times 10^{-12} \quad [1] \text{ mol}^3 \text{ dm}^{-9} \quad [1]$	
(ii)	$K_{sp} = [\text{Ag}^+]^2[\text{CO}_3^{2-}]$ $8.49 \times 10^{-12} = [\text{Ag}^+]^2 \times (2.50 \div (23.0 \times 2 + 12.0 + 16.0 \times 3) \div (100/1000))$ $[\text{Ag}^+] = 6.00 \times 10^{-6} \text{ mol dm}^{-3} \quad [1]$ <p>Solubility of $\text{Ag}_2\text{CO}_3 = \frac{1}{2} \times 6.00 \times 10^{-6} = 3.00 \times 10^{-6} \text{ mol dm}^{-3} \quad [1]$</p>	
(c) (i)	$P_{\text{Al}_2\text{I}_6} = \frac{nRT}{V} = \frac{(\frac{10.0}{816})(8.31)(673)}{2 \times 10^{-3}} = 34300 \text{ Pa} \quad [1]$	
(ii)	<p>$P_{\text{Al}_2\text{I}_6}$ (after dissociation) = $0.65 \times 34300 = 22300 \text{ Pa}$</p> <p>$P_{\text{AlI}_3} = 2 \times 0.35 \times 34300 = 24000 \text{ Pa} \quad [1]$</p> $K_p = \frac{(P_{\text{AlI}_3})^2}{P_{\text{Al}_2\text{I}_6}} = \frac{24000^2}{22300} = 25800 \quad [1] \text{ Pa} \quad [1] \text{ allow ecf}$	
(iii)	<p>ΔS is positive and ΔH is positive. When temperature decreases, <u>$-T\Delta S$ will be smaller than ΔH</u>. [1] <u>ΔG will become more positive and less spontaneous</u>. Hence <u>degree of dissociation decreases</u>. [1]</p> <p>OR</p> <p>When the temperature is decreased, by Le Chatelier's Principle, the <u>backward exothermic reaction is favoured to produce more heat</u>. [1] Hence, the <u>position of equilibrium will shift to the left and degree of dissociation decreases</u>. [1]</p>	

(d) (i)	<p>P:</p>  <p>[1]</p> <p>Q:</p>  <p>[1]</p>
(ii)	Nucleophilic Substitution [1]
(iii)	<p>Step 2: $\text{K}_2\text{Cr}_2\text{O}_7$, $\text{H}_2\text{SO}_4(\text{aq})$, heat [1] [do not accept KMnO_4, $\text{H}_2\text{SO}_4(\text{aq})$]</p> <p>B:</p>  <p>[1]</p> <p>Step 3: HCN, trace amount of $\text{NaOH}(\text{aq})$, $10-20^\circ\text{C}$ [1]</p> <p>C:</p>  <p>[1]</p> <p>Step 4: $\text{H}_2\text{SO}_4(\text{aq})$, heat [1]</p>
3 (a)	<p>Most stable structure:</p> $\text{H}_2\ddot{\text{N}}-\ddot{\text{N}}=\ddot{\text{N}}-\ddot{\text{N}}\text{H}_2 \quad \text{or} \quad \begin{array}{c} \text{H}_2\text{N}-\ddot{\text{N}}-\ddot{\text{N}}=\ddot{\text{N}}-\text{H} \\ \\ \text{H} \end{array}$ <p>[1]</p>
(b)	<p>Both compounds have simple molecular structures. On average, tetrazene is able to form 4 hydrogen bonds per molecule as compared to 1 hydrogen bond per molecule for hydrogen azide. More energy is required to overcome the more extensive hydrogen bonding in tetrazene. Hence, the high boiling point of hydrazine. [1]</p>
(c) (i)	<p>Ammonia is acting as a Bronsted Lowry acid as it donates H^+ to $\text{CH}_2=\text{CH}^-$ to form $\text{CH}_2=\text{CH}_2$. [1]</p>
(ii)	<p>Ammonia is acting as a Lewis base as the N atom donates a lone pair of electrons to the electron deficient B atom. [1]</p>
(iii)	<p>Ammonia is acting as a reducing agent as it reduces the oxidation number of chlorine from +1 in NaOCl to -1 in NaCl. [1]</p>
(iv)	<p>Ammonia is acting as a nucleophile as it contains a lone pair of electrons which attacks the electron deficient carbon of the two acyl chlorides. [1]</p>
(d)	<p>Condensation/ nucleophilic acyl substitution [1]</p>

	<p>[1] for 2 correct sets [2] for all 3 correct sets</p>
(e) (i)	<p>There are 5 electrons in the 2nd shell of the sp² hybridised N.</p> <ul style="list-style-type: none"> • In the three sp² hybrid orbitals of N atom, each hybrid orbital has an unpaired electron. • Sigma bonds are formed when two sp² hybrid orbitals of N atom overlaps head-on with the sp² hybrid orbitals of each carbon atom and when one sp² hybrid orbital of N atom overlaps head on with the 1s orbital of H atom. • The unhybridised p orbital are filled with a lone pair of electrons.
(ii)	<p>4-methylphenylamine is (1) more basic than phthalimide.</p> <p>For 4-methylphenylamine, the (2) p orbital of N atom overlaps with the π electron cloud of the benzene ring, causing the lone pair on N atom to be delocalised into the benzene ring. This (3) decreases the availability of the lone pair of electrons on N atom but lone pair is still available for dative bonding with a H⁺. [1]</p> <p>While in phthalimide, (4) p orbital of N atom overlaps effectively with the π electron cloud of the 2 C=O due to the electronegative O atoms, allowing the lone pair of electrons on N atom to be delocalise into the π electron cloud of C=O, making (5) lone pair not available to form a dative bond with a H⁺. [1]</p>
(f) (i)	<p>$pK_{a1} = 2.2$ $K_{a1} = 10^{-2.20} = 6.31 \times 10^{-3} \text{ mol dm}^{-3}$ [1]</p>
(ii)	<p>When small amount of base is added,</p> <ul style="list-style-type: none"> • Balanced eqn, with correct structural formula of the protonated lysine and product, • full arrow for equation (reaction is a neutralisation reaction) <p>[1] for both points</p> <p>When a <u>small amount</u> of OH⁻ is added, it is removed by <u>the large reservoir</u> of protonated lysine, hence keeping pH <u>fairly constant</u>. [1]</p>
(iii)	<p>Fig 3.2 shows that end point is achieved when 10cm³ of 1.00 mol dm⁻³ NaOH was added to 10cm³ of fully protonated lysine, HA. Hence $[HA] = 1.00 \text{ mol dm}^{-3}$</p> <p>$n(\text{HA}) \text{ at point A} = 1.00 \times \frac{10}{1000} = 1.00 \times 10^{-2} \text{ mol}$</p> <p>At point B (MBC) $n_{\text{HA}} = n_{\text{A}^-}$</p> $= \frac{1.00 \times 10^{-2}}{2}$ $= 5.00 \times 10^{-3} \text{ mol}$

	$n(\text{NaOH}) \text{ added} = 0.0200 \times \frac{10.0}{1000} = 2.00 \times 10^{-4} \text{ mol}$ $\text{HA} + \text{OH}^- \rightarrow \text{A}^- + \text{H}_2\text{O}$ $n_{\text{HA}} \text{ present after NaOH is added} = (5.00 \times 10^{-3}) - (2.00 \times 10^{-4})$ $= 4.80 \times 10^{-3} \text{ mol}$ $n_{\text{A}^-} \text{ present after NaOH is added} = (5.00 \times 10^{-3}) + (2.00 \times 10^{-4})$ $= 5.20 \times 10^{-3} \text{ mol [1]}$ $K_a = \frac{[\text{H}^+][\text{A}^-]}{[\text{HA}]}$ $6.31 \times 10^{-3} = \frac{[\text{H}^+] \times 5.20 \times 10^{-3}}{4.80 \times 10^{-3}}$ $[\text{H}^+] = 5.82 \times 10^{-3} \text{ mol dm}^{-3}$ $\text{pH} = 2.24 \text{ [1]}$ <p style="text-align: center;">OR</p> <p>ICE table for understanding (not necessary to draw for answer)</p> <table><tr><th></th><th>HA (acid)</th><th>+ NaOH</th><th>→</th><th>NaA (salt)</th><th>+ H₂O</th></tr><tr><td>n(Initial)</td><td>5.00×10^{-3}</td><td>2.00×10^{-4}</td><td></td><td>5.00×10^{-3}</td><td>-</td></tr><tr><td>n(Change)</td><td>-2.00×10^{-4}</td><td>-2.00×10^{-4}</td><td></td><td>$+ 2.00 \times 10^{-4}$</td><td>-</td></tr><tr><td>n(Eqm)</td><td>0.0048</td><td>0</td><td></td><td>0.0052</td><td>-</td></tr></table> $\text{pH} = \text{p}K_a + \lg \frac{[\text{salt}]}{[\text{acid}]}$ $= 2.2 + \lg \left(\frac{0.0052/V}{0.0048/V} \right) \text{ [1]}$ $= 2.23$		HA (acid)	+ NaOH	→	NaA (salt)	+ H ₂ O	n(Initial)	5.00×10^{-3}	2.00×10^{-4}		5.00×10^{-3}	-	n(Change)	-2.00×10^{-4}	-2.00×10^{-4}		$+ 2.00 \times 10^{-4}$	-	n(Eqm)	0.0048	0		0.0052	-
	HA (acid)	+ NaOH	→	NaA (salt)	+ H ₂ O																				
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4 (a) (i)	<p>Ionic radius (Pb^{2+}) = 0.120nm Ionic radius (Mg^{2+}) = 0.065nm Ionic radius of the Pb^{2+} is larger than that of Mg^{2+} but ionic charge remains as 2+. Charge density of the Pb^{2+} is lower and its polarising power is also lower. [1]</p> <p>Pb^{2+} is less able to polarise the electron cloud of the CO_3^{2-} ion in PbCO_3. C–O bonds in CO_3^{2-} is polarised and weakened to a smaller extent. More energy is needed to break the C–O bond. PbCO_3 has the higher decomposition temperature. [1]</p>																								
(b) (i)	I: Electrophilic Substitution [1]																								
(ii)	+2 $1s^2 2s^2 2p^6 3s^2 3p^6 3d^6$ [1] for both																								
(iii)	<p>In the presence of ligands, the partially filled 3d orbitals split into two energy levels, with a small energy gap, ΔE. An electron in the lower energy d orbital absorbs energy in the visible spectrum corresponding to ΔE and becomes excited to a vacant d orbital at the higher energy level (d-d transition). [1]</p>																								

	Unabsorbed wavelengths are transmitted and the colour observed is complementary to the colour absorbed. [1]
(iv)	With water as the ligand, electromagnetic waves with energy corresponding to red is absorbed which has a lower energy than blue, which is absorbed when C_5H_5^- is the ligand. [1] C_5H_5^- causes a larger 3d orbital split. [1]
(v)	$E^\circ(\text{Fe}(\text{C}_5\text{H}_5)_2^+ \text{Fe}(\text{C}_5\text{H}_5)_2)$ is less positive than $E^\circ(\text{Fe}^{3+} \text{Fe}^{2+})$. In reaction (II), ferrocene has been oxidised by Fe^{3+} , thus $E^\circ(\text{Fe}(\text{C}_5\text{H}_5)_2^+ \text{Fe}(\text{C}_5\text{H}_5)_2)$ has a less positive E° value than $E^\circ(\text{Fe}^{3+} \text{Fe}^{2+})$ since $E^\circ_{\text{cell}} = E^\circ_{\text{reduced}} - E^\circ_{\text{oxidised}} > 0$. [1]
(vi)	$2\text{Fe}(\text{C}_5\text{H}_5)_2 + \text{S}_2\text{O}_8^{2-} \rightarrow 2\text{Fe}(\text{C}_5\text{H}_5)_2^+ + 2\text{SO}_4^{2-}$ [1] $2\text{Fe}(\text{C}_5\text{H}_5)_2^+ + 2\text{I}^- \rightarrow 2\text{Fe}(\text{C}_5\text{H}_5)_2 + \text{I}_2$ [1] Without the catalyst, the reaction has a high activation energy due to collisions between ions of the same charge. In the presence of catalyst, collisions between ions of the same charge are avoided hence it lowers the activation energy. [1]
(vii)	$+0.54\text{V} < E^\circ(\text{Fe}(\text{C}_5\text{H}_5)_2^+ \text{Fe}(\text{C}_5\text{H}_5)_2) < +0.77\text{V}$ [1]
(c)	<div style="display: flex; justify-content: space-around; align-items: center;">    </div> <p style="text-align: center;">G H J</p> <p>[1] for each correct structure. Maximum [3].</p> <p>G contains an ester and undergoes acid hydrolysis to form a carboxylic acid and an alcohol.</p> <p>G contains an alkene and undergoes oxidative cleavage with acidified MnO_4^- to give H, J and CO_2.</p> <p>G contains two alkenes and undergoes electrophilic addition with two Br_2 to form alkylbromide due to the addition of four Br atoms.</p> <p>J contains a carboxylic acid and undergoes neutralisation with NaHCO_3.</p> <p>J contains a carboxylic acid and undergoes reduction with LiAlH_4 to give a primary alcohol.</p> <p>Primary alcohol cannot undergo elimination with conc H_2SO_4 as the C atom next to C with $-\text{OH}$ group has no available H atom.</p>

	<p>H contains a ketone and undergoes condensation with 2,4-DNPH to give an orange ppt.</p> <p>H contains a ketone since there is no reaction with Tollen's reagent.</p> <p>H contains two $\text{CH}_3\text{CO}-$ and undergoes oxidation with alkaline aqueous iodine to give two moles of yellow ppt of CHI_3. [4]</p>
5 (a) (i)	$\text{BE}(\text{C}-\text{C}) + 2 \times \text{BE}(\text{C}-\text{O}) = 350 + 2 \times 360 = +1070 \text{ kJ mol}^{-1}$ [1]
(ii)	$\Delta H_r^\ominus = \sum \text{BE}(\text{bonds broken}) - \sum \text{BE}(\text{bonds formed})$ $-354 = [4 \times \text{BE}(\text{C}-\text{H}) + \text{BE}(\text{C}=\text{C})] - [\text{BE}(\text{C}-\text{C}) + 4 \times \text{BE}(\text{C}-\text{H}) + 2 \times \text{BE}(\text{C}-\text{O})]$ $-354 = [(4 \times 410) + 610] - [\text{BE}(\text{C}-\text{C}) + (4 \times 410) + 2 \times \text{BE}(\text{C}-\text{O})]$ $-354 = 610 - [\text{BE}(\text{C}-\text{C}) + 2 \times \text{BE}(\text{C}-\text{O})]$ $\text{BE}(\text{C}-\text{C}) + 2 \times \text{BE}(\text{C}-\text{O}) = +964 \text{ kJ mol}^{-1}$ [1] <p>The actual bond energy is less endothermic as it takes lesser energy to break the weaker covalent bonds in the 3 member ring. [1]</p>
(iii)	<p>ΔS is negative. Entropy of the system decreases and there is an decrease in the total number of gaseous particles in the system (i.e. from 2 mol of gaseous reactant to 1 mol of gaseous product). [1] There are less ways of arranging the particles in the system, resulting in a more disordered system to a less disordered system. [1]</p>
(iv)	<p>ΔH^\ominus and ΔS^\ominus are both negative. For reaction to be feasible, $\Delta G < 0$ (negative). Since $\Delta G = \Delta H - T\Delta S$, $T\Delta S$ factor would be smaller in magnitude at lower temperature than ΔH. The reaction is feasible at low temperature. [1]</p>
(b)	 <p>[3] total [1] for each correct step</p>

(c)	<p>Energy /kJ mol⁻¹</p> <p> $\text{Mg(g)} + 2\text{N(g)} + 6\text{O(g)}$ $+148 + 944 + 3(496) = +2580$ $\text{Mg(s)} + \text{N}_2\text{(g)} + 3\text{O}_2\text{(g)}$ -791 $\text{Mg(NO}_3)_2\text{(s)}$ </p> <p> $\text{Mg(g)} + 2\text{NO}_3\text{(g)}$ $+736 + 1450 + 2(-317) = +1552$ $\text{Mg}^{2+}\text{(g)} + 2\text{NO}_3^-\text{(g)}$ </p> <p> $-2[3(400)]$ L.E </p> <p>By Hess' Law: $-791 = (+2580) + (-2400) + (1552) + \text{L.E.}$ $\text{L.E.} = -2523 \text{ kJ mol}^{-1}$ $= -2520 \text{ kJ mol}^{-1} [1]$ </p>
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(d) (i)	2,3-dioxobutanedioic acid								
(ii)	<div style="display: flex; justify-content: space-around; align-items: center;"> <div style="text-align: center;"> J </div> <div style="text-align: center;"> K </div> </div> <p>1 mark for each correct structure</p> <table border="1" style="width: 100%; border-collapse: collapse;"> <thead> <tr> <th style="width: 50%;">Observations</th><th style="width: 50%;">Deductions</th></tr> </thead> <tbody> <tr> <td>J, C₇H₁₀O₂ decolourises bromine water</td><td>J could be alkene or phenol. An <u>alkene</u> undergoes <u>electrophilic addition</u> reaction with Br₂(aq) to form <u>bromohydrin</u>.</td></tr> <tr> <td>22.7 dm³ of hydrogen is formed at s.t.p when J is reacted with sodium metal</td><td><u>2 moles of secondary alcohol in J</u> undergoes <u>reduction</u> with Na to form <u>1 mol of H₂</u>. $n(\text{H}_2) = \frac{22.7}{22.7} = \underline{1 \text{ mol of H}_2}$ is formed. ROH + Na → RO⁻Na⁺ + ½H₂ Since 1 FG produces ½ mole of H₂, there must be <u>2 -OH groups present</u>. -COOH group is absent as it will only produce only ½ mole of H₂ and 2 -COOH groups or 1 -OH & 1 -COOH groups cannot be present as it will not correspond to the molecular formula of J.</td></tr> <tr> <td>On heating with acidified KMnO₄, K, C₅H₄O₅, is the only</td><td><u>Secondary alcohol and the alkene in J</u> undergoes <u>oxidation</u> with KMnO₄ to form <u>ketones</u> and <u>carboxylic acid</u>. There is a decrease in 2C atoms – which suggest that ethane-1,2-dioic acid was oxidised to form CO₂.</td></tr> </tbody> </table>	Observations	Deductions	J , C ₇ H ₁₀ O ₂ decolourises bromine water	J could be alkene or phenol. An <u>alkene</u> undergoes <u>electrophilic addition</u> reaction with Br ₂ (aq) to form <u>bromohydrin</u> .	22.7 dm ³ of hydrogen is formed at s.t.p when J is reacted with sodium metal	<u>2 moles of secondary alcohol in J</u> undergoes <u>reduction</u> with Na to form <u>1 mol of H₂</u> . $n(\text{H}_2) = \frac{22.7}{22.7} = \underline{1 \text{ mol of H}_2}$ is formed. ROH + Na → RO ⁻ Na ⁺ + ½H ₂ Since 1 FG produces ½ mole of H ₂ , there must be <u>2 -OH groups present</u> . -COOH group is absent as it will only produce only ½ mole of H ₂ and 2 -COOH groups or 1 -OH & 1 -COOH groups cannot be present as it will not correspond to the molecular formula of J .	On heating with acidified KMnO ₄ , K , C ₅ H ₄ O ₅ , is the only	<u>Secondary alcohol and the alkene in J</u> undergoes <u>oxidation</u> with KMnO ₄ to form <u>ketones</u> and <u>carboxylic acid</u> . There is a decrease in 2C atoms – which suggest that ethane-1,2-dioic acid was oxidised to form CO ₂ .
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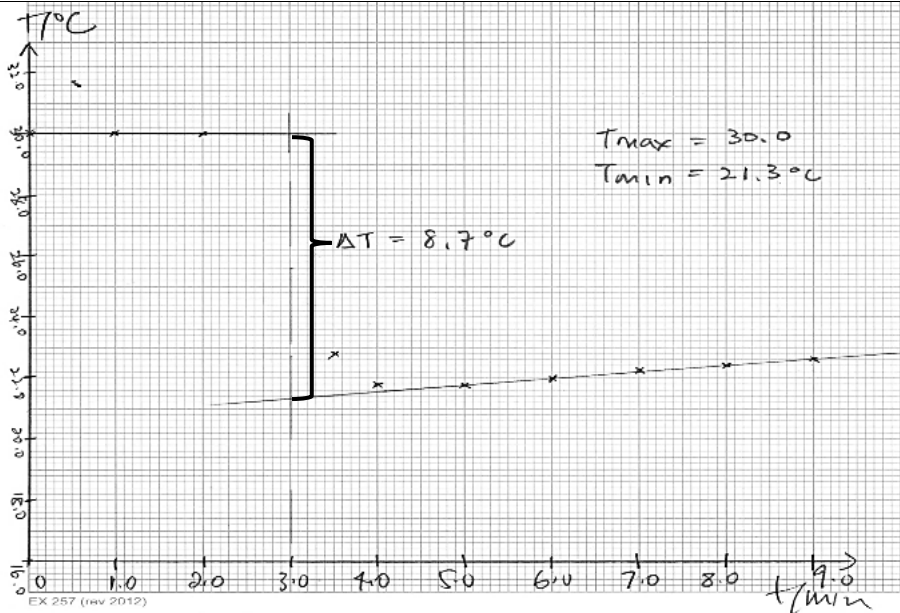
	organic product formed	From the given structure of L , it can be seen that K is unlikely a phenol and should not contain benzene as there will not be oxidation that leads to a reduction of 2 C atoms.
	K reacts with sodium carbonate and 2,4 DNPH	<u>Carboxylic acid</u> in K undergoes <u>acid-base</u> with <u>Na₂CO₃</u> to form <u>salt</u> and water. <u>Ketone</u> in K undergoes <u>condensation</u> with <u>2,4 DNPH</u> to give an <u>orange ppt</u> of hydrazone.
	K reacts with alkaline aqueous iodine to form L	<u>Methyl ketone</u> in K undergoes <u>oxidation</u> with <u>alkaline I₂(aq)</u> to give a carboxylate salt and <u>yellow ppt</u> of CHI ₃ . The <u>carboxylate ion</u> undergoes <u>acid-base reaction</u> with acid to form <u>COOH</u> in L .
[4]		

Paper 4 Answers

1	(a)	$\dots 2 \text{MnO}_4^- + \dots 16 \text{H}^+ + \dots 5 \text{C}_2\text{O}_4^{2-} \rightarrow \dots 2 \text{Mn}^{2+} + \dots 8 \text{H}_2\text{O} + \dots 10 \text{CO}_2$ [1]																				
	(b)	Dilution of FA 2 <table><tr><td>final burette reading / cm³</td><td colspan="2">42.50</td></tr><tr><td>initial burette reading / cm³</td><td colspan="2">0.00</td></tr><tr><td>volume of FA 2 added / cm³</td><td colspan="2">42.50</td></tr></table> Titration of FA 3 <table><tr><td>final burette reading / cm³</td><td>26.15</td><td>26.15</td></tr><tr><td>initial burette reading / cm³</td><td>0.00</td><td>0.00</td></tr><tr><td>volume of FA 1 added / cm³</td><td>26.15</td><td>26.15</td></tr></table> <u>For both tables:</u> <ul style="list-style-type: none">correct heading with units [1]data recorded to 2 d.p. [1]			final burette reading / cm ³	42.50		initial burette reading / cm ³	0.00		volume of FA 2 added / cm ³	42.50		final burette reading / cm ³	26.15	26.15	initial burette reading / cm ³	0.00	0.00	volume of FA 1 added / cm ³	26.15	26.15
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	(c)	Ave $V_{\text{FA1}} = 26.15 \text{ cm}^3$ <ul style="list-style-type: none">Correctly calculated ave (V_{FA1}) and consistent choices [1]Scaled titre diff ≤ 0.3 [2]; Scaled titre diff ≤ 0.5 [1]																				
	(d)	(i)	$n_{\text{MnO}_4^-}$ present in $V_{\text{FA 1}}$ $= 0.023 \times \frac{26.15}{1000}$ $= 6.014 \times 10^{-4} \text{ mol}$																			
		(ii)	$n_{\text{C}_2\text{O}_4^{2-}}$ that reacted with MnO_4^{2-} $= 6.014 \times 10^{-4} \times \frac{5}{2}$ ecf ratio from 1(a) $= 1.503 \times 10^{-3} \text{ mol}$																			
		(iii)	$[\text{C}_2\text{O}_4^{2-}]$ in FA 3 $= 0.0000575 V_{\text{FA 1}} \div \frac{25.0}{1000}$ $= 0.0023 V_{\text{FA 1}} \text{ mol dm}^{-3}$																			
		(iv)	$[\text{C}_2\text{O}_4^{2-}]$ in FA 2 $\times \frac{42.5}{1000} = 6.014 \times 10^{-2} \times \frac{250}{1000}$ $[\text{C}_2\text{O}_4^{2-}]$ in FA 2 $= 6.014 \times 10^{-2} \times \frac{250}{1000} \times \frac{1000}{42.5}$ $= 0.3537 \approx 0.354 \text{ mol dm}^{-3}$																			
		(v)	M_r of $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O} = 126.0$ $[\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}]$ in g dm^{-3} $= 0.3537 \times 126.0$ $= 44.57 \approx 44.6 \text{ g dm}^{-3}$																			

(e) (i)	Initial reaction is <u>between 2 negatively charged</u> MnO_4^- and $\text{C}_2\text{O}_4^{2-}$ <u>ions</u> that <u>repels</u> (or <u>has high activation energy</u>) and requires longer time to react. [1]
(ii)	Mn^{2+} <u>ions</u> produced in the initial reaction <u>catalyse subsequent reactions</u> and increases the rate of reaction. [1]

2 (a)	<table><tr><td>Mass of bottle and solid FA 4 / g</td><td></td></tr><tr><td>Mass of bottle and residual FA 4 / g</td><td></td></tr><tr><td>Mass of FA 4 used / g</td><td>7.85</td></tr></table> <table><tr><td>t / min</td><td>0.0</td><td>1.0</td><td>2.0</td><td>3.5</td><td>4.0</td><td>5.0</td><td>6.0</td><td>7.0</td><td>8.0</td><td>9.0</td></tr><tr><td>Temp / °C</td><td>30.0</td><td>30.0</td><td>30.0</td><td>22.8</td><td>21.8</td><td>21.8</td><td>22.0</td><td>22.2</td><td>22.4</td><td>22.6</td></tr></table> <p>[1] correct header + units [1] mass recorded to 2dp/3dp [1] temp recorded to 1dp, time to nearest 0.5 min [1] no temp recorded at t = 3.0 min, temp, T is recorded every min</p>	Mass of bottle and solid FA 4 / g		Mass of bottle and residual FA 4 / g		Mass of FA 4 used / g	7.85	t / min	0.0	1.0	2.0	3.5	4.0	5.0	6.0	7.0	8.0	9.0	Temp / °C	30.0	30.0	30.0	22.8	21.8	21.8	22.0	22.2	22.4	22.6
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(b)	 <p>[1] scale – no odd scale and occupy more than half the plot [1] extrapolation both best-fit straight line to t = 3.0 min [1] all data points are plotted and two best-fit straight lines drawn [1] second best-fit straight line follows an increasing trend</p>
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(c)	$T_{\min} = 21.3\text{ }^{\circ}\text{C}$ $T_{\max} = 30.0\text{ }^{\circ}\text{C}$ $\Delta T = 30.0 - 21.3 = 8.7\text{ }^{\circ}\text{C}$ [1]
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(d)	Heat change, $q = 50 \times 1.00 \times 4.18 \times 8.7$ $= 1818 \approx 1820\text{ J}$ [1]
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(e)	$\Delta H_{\text{sol}} = + \frac{1818}{7.85 / 74.6 \times 1000}$ $= +17.27 \approx +17.3 \text{ kJ mol}^{-1}$ <ul style="list-style-type: none"> Working for Q/n & /1000 (+) sign for ΔH as the reaction is endothermic Correct calculation of answer (3 sf answer)
(f)	<p>Since the same mass of solid FA 4 is used, heat absorbed by reaction, q, would remain the same.</p> <p>But volume of water is doubled, so ΔT would be halved. [1]</p> <p>$\Delta H = q/n$ Since q, and amount of solid used remained the same, ΔH_{sol} would remain the same. [1]</p>
3 (a) (i)	<p>NaHCO_3 reacts with the H_2SO_4 catalyst, and slows down/stops the reaction.</p>
(ii)	<p>NaHCO_3 will effervesce when reacting with the acid, and when there is no more effervescence, this gives an indication that the reaction has stopped.</p> <p>OR</p> <p>NaOH reacts with the reactants propanone and iodine.</p>
(b)	<ol style="list-style-type: none"> Using a 50 cm³ measuring cylinder, add 30 cm³ of CH_3COCH_3 to a 250 cm³ conical flask. Using another 50 cm³ measuring cylinder, add 30 cm³ of H_2SO_4 to the same conical flask. Using another 50 cm³ measuring cylinder, add 40 cm³ of I_2 to the same conical flask and immediately swirl the mixture thoroughly to mix its contents. Start the stopwatch. Steps 1 - 3: [1] apparatus, correct volumes of reactants [1] start timing when either I_2 or CH_3COCH_3 is added last Using a 25 cm³ measuring cylinder, measure 25 cm³ of $\text{NaHCO}_3(\text{aq})$. Using a 10.0 cm³ pipette, pipette 10.0 cm³ of the reaction mixture into a 100 cm³ conical flask. Add in the 25 cm³ of $\text{NaHCO}_3(\text{aq})$ to the 10.0 cm³ reaction mixture. Steps 4 - 5: [1] withdrawing 10 cm³ samples with pipette [1] quenching with 10 – 25 cm³ NaHCO_3. Immediately titrate the I_2 in the 100 cm³ conical flask with $\text{Na}_2\text{S}_2\text{O}_3$ until the solution turns pale yellow. Add in 10 drops of starch indicator and continue titrating till solution turns from dark blue to colourless. Record your titration results. [1] for stating titration (do not penalise if missing details like addition of starch indicator/end point colour change) Rinse the 100 cm³ conical flask with water. Repeat steps 5 – 7, five more times, when the stop-watch shows 2, 5, 10, 15 and 20 minutes. Step 8: [1] for repeating 5 or 4 more times at suitable timed intervals. (Award mark as long as student repeats experiment at least 5 times ie minimum 5 data points) (Do not penalise for duration of experiment i.e. experiment can be more than 20min although this is not recommended for continuous method) Plot a graph of volume ($\text{Na}_2\text{S}_2\text{O}_3$) against time.

(c)

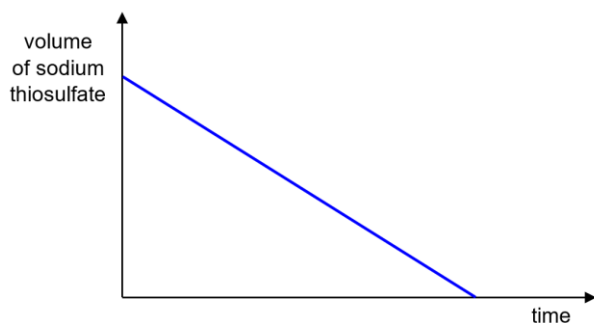


Fig. 3.1

[1] for straight line graph

The graph is a straight line with a negative constant gradient.

Since $V(\text{Na}_2\text{S}_2\text{O}_3) \propto [\text{I}_2]$, and $[\text{I}_2]$ decreases linearly with time, rate is independent of $[\text{I}_2]$.

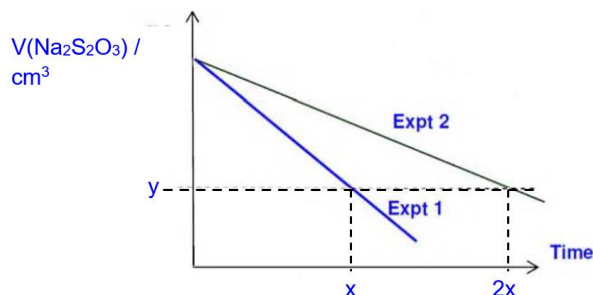
[1] Hence order w.r.t. $[\text{I}_2]$ is zero.

(d)

Experiment	Aim	$V(\text{I}_2)$ /cm ³	$V(\text{CH}_3\text{COCH}_3)$ / cm ³	$V(\text{H}_2\text{SO}_4)$ / cm ³	$V(\text{water})$ /cm ³
1	Determine order w.r.t. I_2	40	30	30	0
2	Determine order w.r.t. CH_3COCH_3	40	15	30	15

[1] for varying volume of propanone and keeping total volume of mixture constant.

Accept variations of volume of propanone (like 10cm³), but answer in explanation should correspond accordingly (gradient for graph for experiment 2 should be 1/3 that of experiment 1)



[1] for sketch of graph for expt 2 (less steep than expt 1) on Fig 3.1 in (c)

- $[\text{CH}_3\text{COCH}_3] \propto V(\text{CH}_3\text{COCH}_3)$ since total volume of mixture was kept constant.
- The gradient of the curves gives the rate.
- Comparing experiments 1 and 2, as $V(\text{CH}_3\text{COCH}_3)$ increases by 2 times, the initial rate should also increase approximately 2 times ($\frac{y/x}{y/2x}$).
- Hence rate is **directly proportional** to $[\text{CH}_3\text{COCH}_3]$, order of reaction w.r.t CH_3COCH_3 is 1.

[1] for 1 or 2 points, [2] for 4 points [1] for 1 or 2 points, [2] for 4 points

4 (a) (i)	tests		observations																									
	1	To 3 cm depth of dilute nitric acid in a boiling tube, add half the amount of FA 5 from the bottle. Use the resulting solution, labelled FA 6 , for tests 3–7.	<ul style="list-style-type: none">• Effervescence (observed). Gas evolved gives white ppt with lime water.• Yellow solution formed. [1]																									
	2	To 1 cm depth of FA 6 , add 1 cm depth of sodium chlorate(I). To the resulting solution, add aqueous sodium thiosulfate dropwise.	<ul style="list-style-type: none">• Brown solution formed. Purple vapour evolved.• Brown solution decolourised. [1]																									
	3	To 1 cm depth of FA 6 , add aqueous sodium hydroxide.	<ul style="list-style-type: none">• White ppt formed soluble in excess aq NaOH to give a colourless solution. [1]																									
	4	To 1 cm depth of FA 6 , add aqueous ammonia.	<ul style="list-style-type: none">• White ppt formed soluble in excess aq NH₃ to give a colourless solution. [1]																									
	5	To 1 cm depth of FA 6 , add 1 cm depth of copper(II) sulfate. To the resulting solution, add aqueous sodium thiosulfate dropwise until no further change occurs.	<ul style="list-style-type: none">• Cream/light brown ppt formed in brown solution.• Brown solution decolourised leaving white ppt. (White ppt dissolved in excess Na₂S₂O₃) [1]																									
	6	To 1 cm depth of FA 6 , add a few drops of aqueous silver nitrate. To the resulting solution, add aqueous ammonia until no further change occurs.	<ul style="list-style-type: none">• Yellow ppt formed.• Yellow ppt insoluble in excess aq NH₃. [1]																									
(ii)		Ions present: <u>Zn²⁺, CO₃²⁻, I⁻</u> [1]																										
(b) (i)		<table><tr><td colspan="2" rowspan="2"></td><td colspan="3">Observations</td></tr><tr><td>FA 7</td><td>FA 8</td><td>FA 9</td></tr><tr><td>1</td><td>2,4-DNPH</td><td>Orange ppt formed</td><td>Orange ppt formed</td><td>no observable change</td></tr><tr><td>2</td><td>Tollens' reagent</td><td>No silver mirror</td><td>Silver mirror formed</td><td>No silver mirror</td></tr><tr><td>3</td><td>Sodium carbonate powder</td><td>no observable change</td><td>no observable change</td><td>Effervescence. Gas evolved gives white ppt with lime water.</td></tr></table>						Observations			FA 7	FA 8	FA 9	1	2,4-DNPH	Orange ppt formed	Orange ppt formed	no observable change	2	Tollens' reagent	No silver mirror	Silver mirror formed	No silver mirror	3	Sodium carbonate powder	no observable change	no observable change	Effervescence. Gas evolved gives white ppt with lime water.
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(ii)		FA 7: Propanone FA 8: Propanal FA 9: Propanoic acid [1]																										

(iii)	Reagents and conditions: $\text{I}_2(\text{aq})$, $\text{NaOH}(\text{aq})$, warm [1] Observation: yellow ppt of CHI_3 formed [1]
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